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diagrams are one of the most important sources of information concerning the behavior of elements, compounds and solutions. They provide us with the knowledge of phase composition and phase stability as a function of temperature (T), pressure (P) and composition (C).

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Fluid Phase Equilibria - Journal - Elsevier  
Phase Equilibria and Salt Effect on the Aqueous Two-Phase System of Polyoxyethylene Cetyl Ether and Sulfate Salt at Three Temperatures. Journal of Chemical & Engineering Data 2016 , 61 (6) , 2135-2143.

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Measurement and Correlation of Phase Equilibria in Aqueous

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Phase Equilibria in the  $H_2/C_2H_4$  System at Temperatures from 114.1 to 247.1 K and Pressures to 600 MPa Andreas Heintz. School of Chemical Engineering, Cornell University, Ithaca, NY 14853, USA. Search for more papers by this author. William B. Streett.

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The phase rule is a general principle governing "pVT systems" in thermodynamic equilibrium, whose states are completely described by the variables pressure (p), volume (V) and temperature (T). If F is the number of degrees of

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freedom, C is the number of components and P is the number of phases, then 
$$F=C-P+2.$$

Phase rule - Wikipedia

Phase Equilibria in Hydrocarbon - Water Systems

(Department of Chemical Engineering, The Pennsylvania State University, Report No. API-7-77). Kabadi, Vinayak and Ronald P. Danner and The Department of Chemical Engineering, The Pennsylvania State University. Published by Department of Chemical Engineering, The Pennsylvania State University. (1977)

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